Surname	Centre Number	Candidate Number
First name(s)		2

GCE AS

B410U20-1

O20-B410U20-1

Part of WJEC

TUESDAY, 13 OCTOBER 2020 – MORNING

CHEMISTRY – AS component 2

Energy, Rate and Chemistry of Carbon Compounds

1 hour 30 minutes

	For Examiner's use only		
	Question	**Maximum Mark**	**Mark Awarded**
Section A	**1. to 6.**	**10**	
Section B	**7.**	**13**	
	8.	**19**	
	9.	**11**	
	10.	**13**	
	11.	**14**	
	Total	**80**	

B410U201
01

ADDITIONAL MATERIALS

In addition to this examination paper, you will need a:
- calculator;
- **Data Booklet** supplied by WJEC.

INSTRUCTIONS TO CANDIDATES

Use black ink or black ball-point pen.

Write your name, centre number and candidate number in the spaces at the top of this page.

Section A Answer **all** questions in the spaces provided.
Section B Answer **all** questions in the spaces provided.

Candidates are advised to allocate their time appropriately between **Section A (10 marks)** and **Section B (70 marks)**.

INFORMATION FOR CANDIDATES

The number of marks is given in brackets at the end of each question or part-question.

The maximum mark for this paper is 80.

Your answers must be relevant and must make full use of the information given to be awarded full marks for a question.

The assessment of the quality of extended response (QER) will take place in **Q.10***(a)*(i).

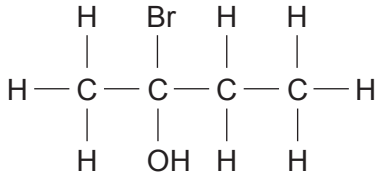
If you run out of space, use the continuation page(s) at the back of the booklet, taking care to number the question(s) correctly.

Examiner only

SECTION A

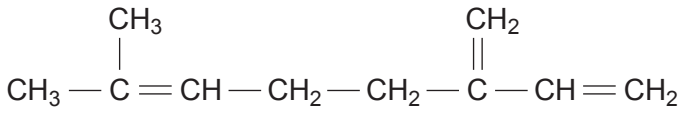
Answer **all** *questions in the spaces provided.*

1. Name the compound whose formula is shown below. [1]



.....

2. Myrcene is a significant component of the essential oils of many plants. It has the structure shown below.



(a) State the **empirical** formula of myrcene. [1]

.....

(b) Draw the **skeletal** formula of myrcene. [1]

(c) Describe a test to show that myrcene contains C=C double bonds. [2]

.....

.....

Examiner only

3. Name an element that contaminates many hydrocarbon fuels and that burns to form an oxide that contributes to acid rain. [1]

.....

4. The polymer PMMA which is used in aeroplane windows and skylights is formed from the monomer methyl methacrylate.

$CH_2{=}C(CH_3)CO_2CH_3$

Draw the repeat unit of this polymer. [1]

5. Complete the equation for the reaction between propanoic acid and sodium carbonate. [1]

$2C_2H_5COOH + Na_2CO_3 \longrightarrow$ .....

6. Collagen has the formula $C_{57}H_{91}N_{19}O_{16}$. Analysis of a sample showed that it contained 0.0204 mol of carbon.

Calculate the simplest ratio of carbon atoms to nitrogen atoms in the formula and use this to calculate the mass of nitrogen present in the sample. [2]

Mass of nitrogen = ... g

Examiner only

SECTION B

Answer **all** *questions in the spaces provided.*

7. *(a)* Biogas refers to a mixture of different gases produced by the breakdown of organic matter in the absence of oxygen. Its main component is methane.

Suggest why burning biogas is considered to be more environmentally friendly than burning natural gas. [2]

.....

.....

.....

.....

Examiner only

(b) Ethanol is widely used as a biofuel in some countries.

(i) Write the equation to represent the standard molar enthalpy change of formation of ethanol. [2]

.....

(ii) The equation for its combustion is given below.

$$C_2H_5OH + 3O_2 \longrightarrow 2CO_2 + 3H_2O$$

I. A student calculated the enthalpy change for this reaction as –1031 kJ mol^{-1} using the average bond enthalpy values given in the table.

Calculate the average bond enthalpy for the C — O bond. [3]

Bond	Average bond enthalpy / kJ mol^{-1}
C — C	348
C — H	412
O — H	463
O = O	496
C = O	743

Average bond enthalpy of C — O = ... kJ mol^{-1}

II. The literature value for the enthalpy change of this reaction is –1370 kJ mol^{-1}.

State why the values are different. [1]

.....

.....

Examiner only

(c) Ethanol is the main alcohol present in alcoholic drinks. When driving in England, the legal limit for alcohol in the blood is 80 mg per 100 cm^3 of blood.

Calculate the concentration of ethanol in blood at this level in mol dm^{-3}. [2]

Concentration = ... mol dm^{-3}

(d) A fluorocarbon has a relative molecular mass which is twice that of its empirical formula. 9.56 g of this compound contains 6.73 g of fluorine.

Find its molecular formula. [3]

Molecular formula .....

13

BLANK PAGE

Examiner only

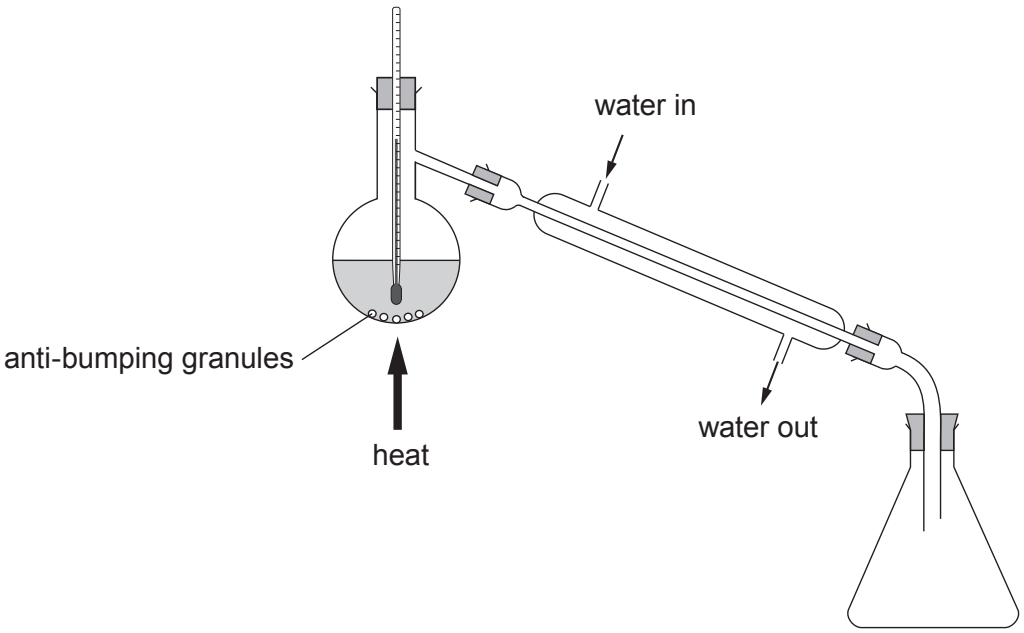
8. *(a)* A student wanted to form a halogenoalkane from an alcohol. He added hydrochloric acid to the alcohol and a mixture of organic and inorganic products formed. The halogenoalkane is the only substance in the mixture that does not dissolve in the inorganic aqueous solution.

(i) Suggest how the student could separate the halogenoalkane from the other compounds. [1]

.....

.....

(ii) The student wants to purify the halogenoalkane by distillation. He sets up the apparatus as shown in the diagram. You may assume that all the equipment is suitably clamped.



I. Anti-bumping granules were placed in the flask.

Suggest why these granules prevent bumping. [1]

.....

.....

Examiner only

II. State **two** changes that must be made to the apparatus for safe and effective use. Give your reason in each case. [4]

.....

.....

.....

.....

.....

.....

(b) Another student wanted to make 1-chloropentane.

She started with pentan-1-ol and obtained 1.62 g of 1-chloropentane. The percentage yield of 1-chloropentane was 67 %.

A fellow student told her that since 67 % is about two-thirds, she must have started with about 2.43 g of pentan-1-ol.

Is the student correct? Justify your answer. [3]

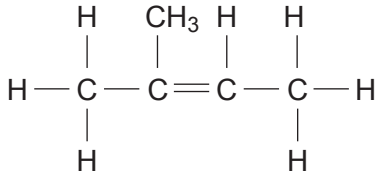
.....

.....

Examiner only

(c) Halogenoalkanes can also be formed from alkenes.

The alkene 2-methylbut-2-ene reacts with hydrogen bromide to form a mixture of 2-bromo-2-methylbutane and 2-bromo-3-methylbutane.



(i) Classify the reaction mechanism. [1]

.....

(ii) Explain why 2-bromo-2-methylbutane is the major product of this reaction. [2]

.....

.....

.....

.....

(iii) Draw the mechanism for the formation of 2-bromo-2-methylbutane. [3]

Examiner only

(d) When 2-bromobutane is heated with potassium hydroxide dissolved in ethanol, two structural isomers are formed.

(i) State the meaning of the term *structural isomers*. [1]

.....

.....

(ii) Draw the structure of both isomers. [2]

(iii) Circle the isomer which exhibits *E*-*Z* isomerism. Explain your choice. [1]

.....

.....

19

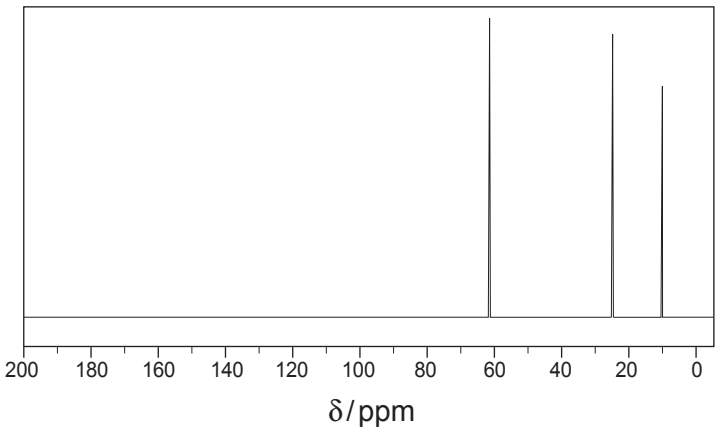
Examiner only

9. *(a)* Compound **X** is an alcohol. When compound **X** is warmed with acidified potassium dichromate(VI) there is a colour change.

In the mass spectrum of compound **X**

- the molecular ion peak is at m/z 74
- the peak with the greatest relative intensity is at m/z 43 and is due to a secondary carbocation

Compound **X** has the following ^{13}C NMR spectrum.



Explain what can be deduced from each piece of information and identify the structure of compound **X**. [7]

.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

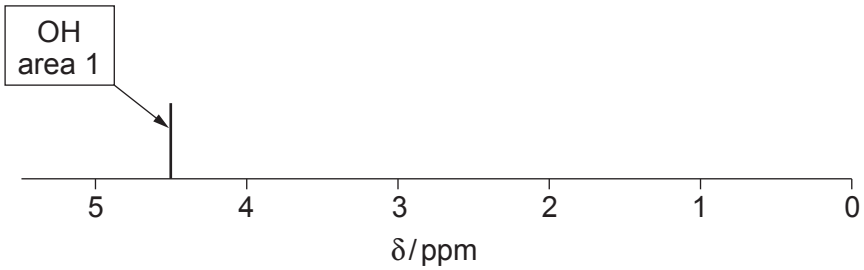
Structure of compound **X**

Examiner only

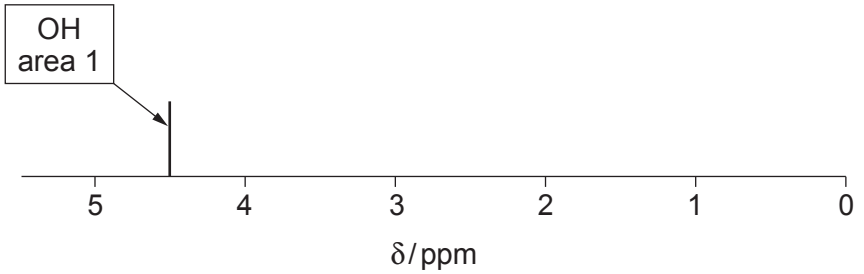
(b) Complete the sketches of the **low resolution** ^{1}H NMR spectra for propan-1-ol and propan-2-ol.

Identify which protons are responsible for each peak giving the approximate chemical shift (ppm) and the relative area of each peak. [4]

Propan-1-ol, $CH_3CH_2CH_2OH$



Propan-2-ol, $CH_3CH(OH)CH_3$



11

Examiner only

10. *(a)* (i) Describe how you would carry out an experiment to determine the enthalpy change of the reaction between magnesium oxide and hydrochloric acid.

Your answer should include details of the apparatus to be used, the measurements to be taken and the way in which you would use your results to determine the enthalpy change. [6 QER]

Examiner only

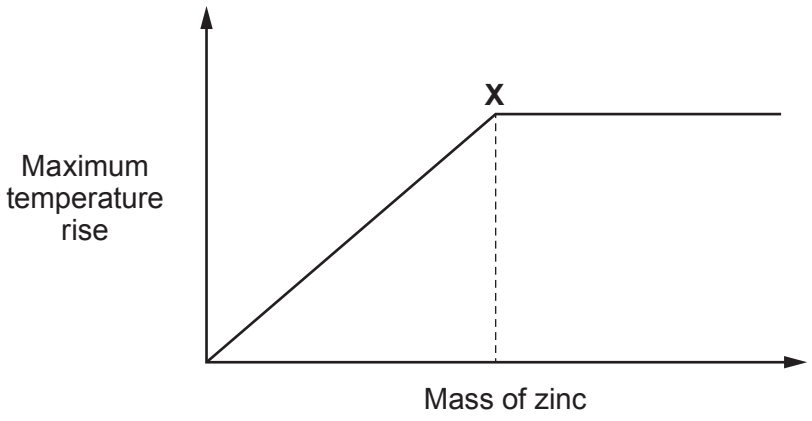
(ii) A student carried out this experiment and calculated the enthalpy change of reaction to be –115 kJ mol^{-1}.

He used 0.900 g of magnesium oxide and 50.0 cm^3 of 2.00 mol dm^{-3} hydrochloric acid.

Show that the acid is in excess and hence calculate the maximum temperature rise the student recorded. [4]

Maximum temperature rise = ... °C

(b) The graph below shows the maximum temperature rise as different masses of zinc react with separate samples of 50.0 cm^3 of 0.100 mol dm^{-3} hydrochloric acid.



$$Zn(s) + 2HCl(aq) \longrightarrow ZnCl_2(aq) + H_2(g)$$

Calculate the mass of zinc at the position labelled **X**. [3]

Mass of zinc = ... g

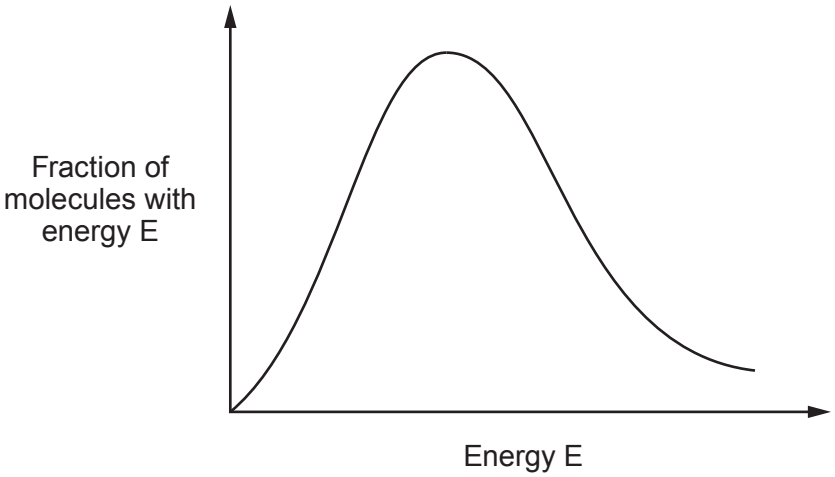
Examiner only

11. *(a)* The decomposition of gaseous hydrogen iodide, HI, is represented by the following equation.

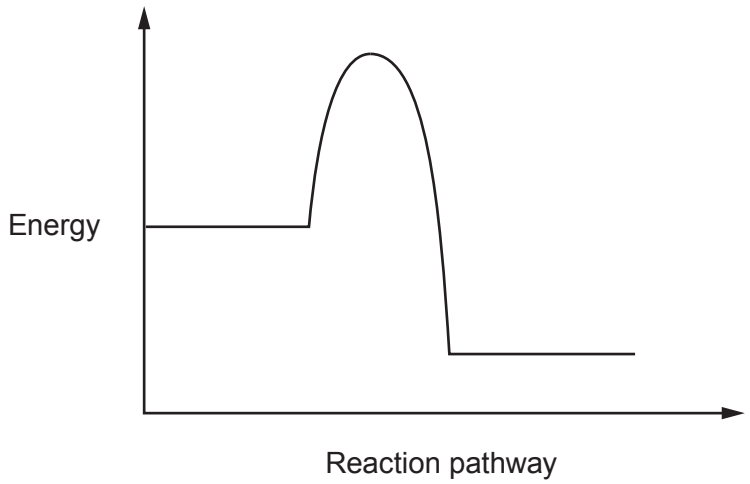
$$2HI(g) \rightleftharpoons H_2(g) + I_2(g) \qquad \Delta H = -53\,kJ\,mol^{-1}$$

(i) Its energy distribution curve at a certain temperature is shown below.

On the same axes, draw another curve to show the distribution at a higher temperature. [1]



(ii) The energy profile for this reaction is shown below.



I. Label the position that represents the transition state of the reaction. [1]

II. **On the same axes**, draw the energy profile for the same reaction if it were catalysed. [1]

Examiner only

(iii) The activation energy of the forward reaction (E_f) is 195 kJ mol^{-1}.

Calculate the activation energy of the reverse reaction (E_b). [1]

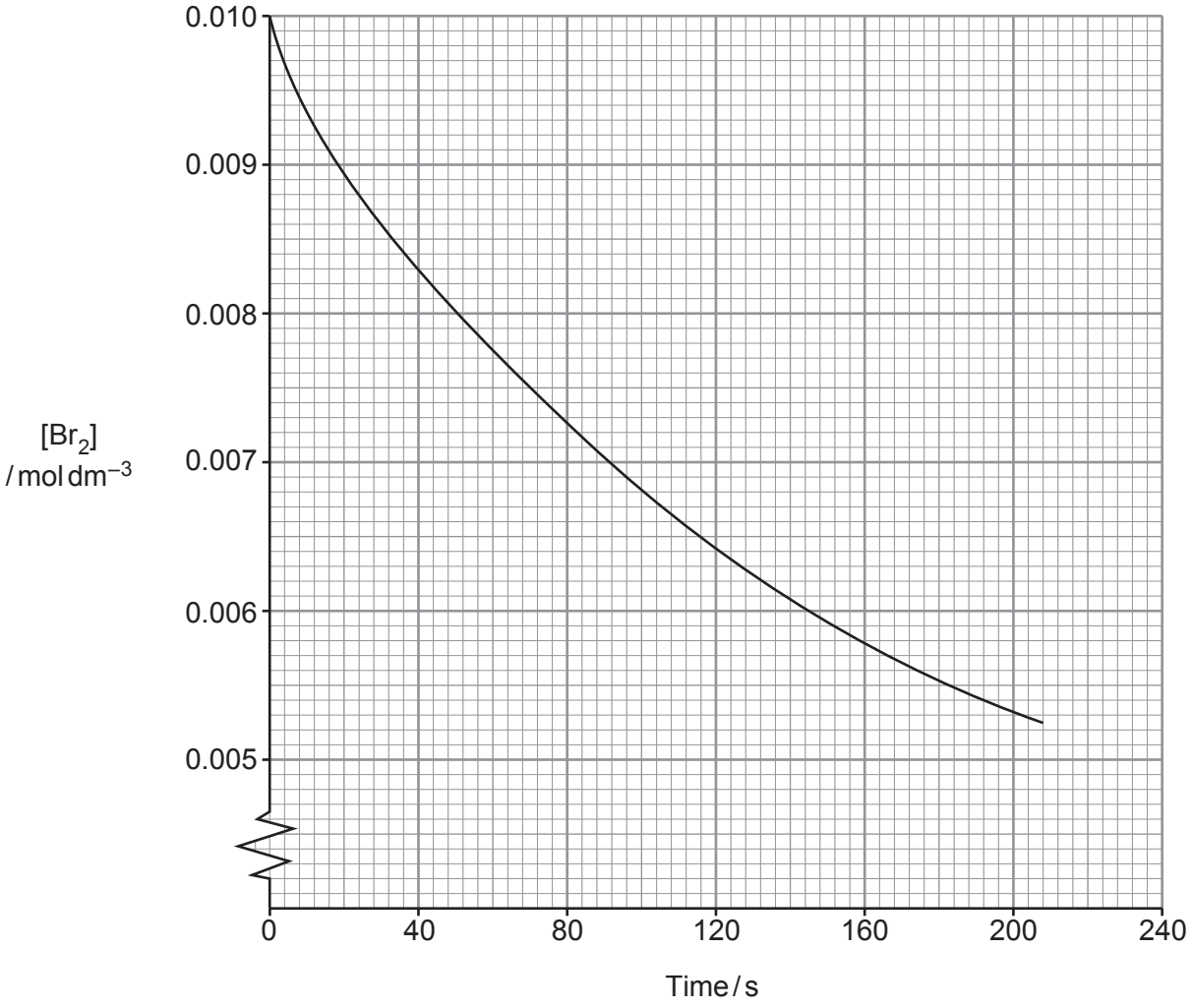
E_b = ... kJ mol^{-1}

Examiner only

(b) Bromine oxidises methanoic acid according to the following equation.

$$Br_2(aq) + HCOOH(aq) \longrightarrow 2H^+(aq) + 2Br^-(aq) + CO_2(g)$$

The graph below shows how the concentration of bromine changes in the initial stages of the reaction.



(i) Calculate the initial rate of the reaction from the graph and give its unit. **Show your working**. [4]

Initial rate = ...

Unit .....

Examiner only

(ii) Suggest **two** ways in which the rate of this reaction could be measured. [2]

.....

.....

.....

.....

(iii) I. State how the graph shows that the rate decreases as the reaction proceeds. [1]

.....

II. Use collision theory to explain why the rate of the reaction decreases as the reaction proceeds. [2]

.....

.....

.....

.....

(iv) State why it is necessary to keep the temperature constant during this experiment. [1]

.....

.....

14

END OF PAPER

Examiner only

For continuation only.